10.2 Organic Halides

Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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| **Name** | **Formula** | **Line structure** | **Name** | **Formula** | **Line structure** |
| bromomethane |  |  | 2-fluoropropane |  |  |
| chloromethane |  |  | 2,2-dibromopropane |  |  |
| diiodomethane |  |  | hexachlorobenzene |  |  |
| difluoromethane |  |  | meta-difluorobenzene |  |  |
| dichlorodifluoromethane |  |  | 1,1-dichlorocyclopropane |  |  |
| 1,1-dichloroethane |  |  | 1,3-difluoropropane |  |  |
| 1-fluoropropane |  |  | Tetra chloroethene |  |  |

1. Molecules with the same chemical formula but different molecular structures are called as **structural isomers**. Draw/name **four** isomers of dibromopropane.
2. A **chiral** carbon is a carbon centre that has **four different substituents**. A molecule with one chiral centre is said to be optically active (solutions of the molecule can rotate plane-polarized light). The mirror image of a compound with a chiral centre is said to be non-super-imposable. Draw the two mirror image isomers of bromochlorofluoromethane.