5.2 Dissociation Equations

Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Identify which of the following classes of compounds are electrolytes:

(a) soluble ionic compounds

(b) insoluble ionic compounds

(c) soluble molecular compounds

(d) insoluble molecular compounds

(e) molecular compounds that ionize in water (acids)

2. Write dissociation equations for the following compounds in water.

(a) LiF(s)

(b) Ca(CH3COO)2(s)

(c) CH3OH(ℓ)

(d) HgBr(s)

(e) PbSO4(s)

(f) HBr(g) (acid)

(g) CCl4(ℓ)

(h) O2(g)

(i) (NH4)3PO4(s)

3. Identify which of the following compounds are soluble in water.

(a) NaCl(s)

(b) AgCl(s)

(c) BaSO4(s)

(d) C12H22O11(s) (table sugar)

(e) NH4OH(s)

(f) Ca(NO3)2(s)

(g) C25H52(s) (candle wax)

(h) Mg(OH)2(s)